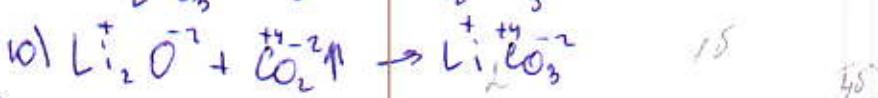
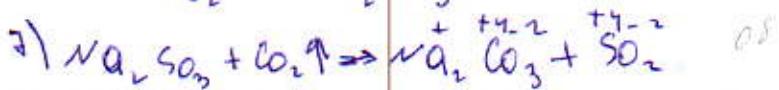
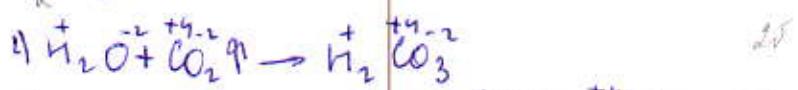
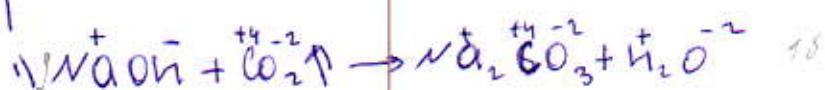
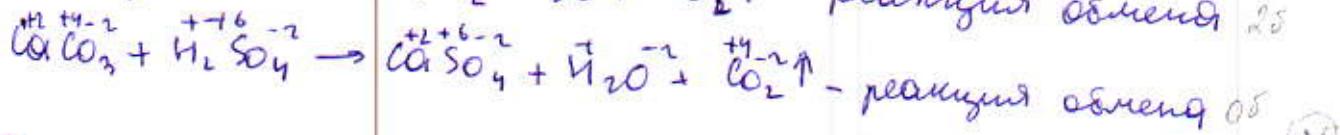
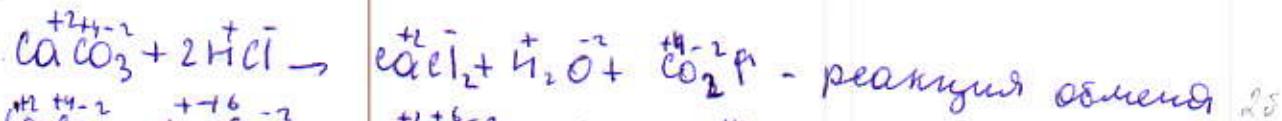


X-9.1

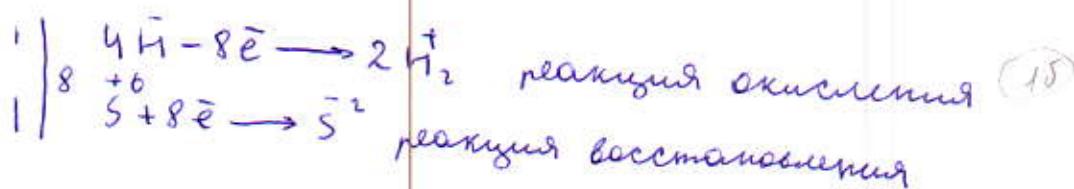
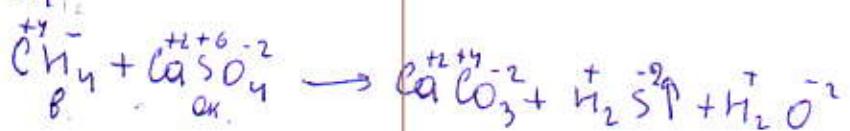
1)



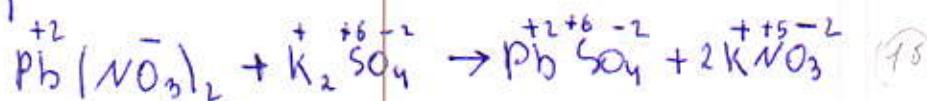
2)



3)



4)



Учеба: 85

X - 9. 2.

№ 1.

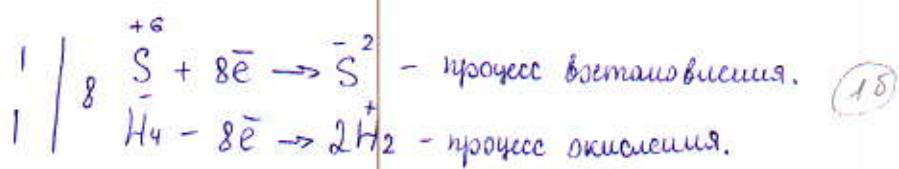
- 1)  $2\text{NaOH} + \text{CO}_2 \rightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O}$  25  
 2)  $\text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3$  25  
 3)  $\text{Na}_2\text{SO}_3 + 2\text{CO}_2 \rightarrow 2\text{Na}_2\text{CO}_3 + 2\text{SO}_2$  05  
 10)  $\text{CO}_2 + \text{Li}_2\text{O} \rightarrow \text{Li}_2\text{CO}_3$  25

(65)

№ 2:

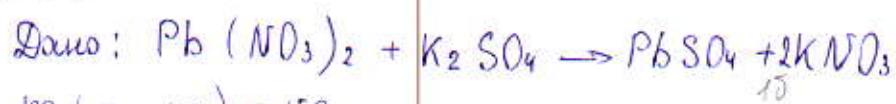
- 1)  $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{CO}_3 \xrightarrow{\text{CO}_2 \uparrow} \text{H}_2\text{O}$  25  
 2)  $\text{CaCO}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{CaSO}_4 + \text{H}_2\text{CO}_3 \xrightarrow[\text{H}_2\text{O}]{\text{CO}_2 \uparrow}$  05

(25)



(18)

№ 4.



$$m(\text{п-ра}) = 150\text{г}$$

$$m(\text{б-ра исчезну}) = 3,31\text{ г}$$

Решение:

$$M(\text{PbSO}_4) = 207 + 16 \cdot 4 + 32 = 303$$

$$n(\text{PbSO}_4) = \frac{m}{M} = \frac{3,31}{303} = 0,01\text{ моль}$$

$$15 n(\text{PbSO}_4) = n(\text{Pb}(\text{NO}_3)_2) = 0,01\text{ моль}$$

$$15 m(\text{Pb}(\text{NO}_3)_2) = n \cdot M = 331 \cdot 0,01 \text{ моль} = 3,31 \text{ г}$$

$$M(\text{Pb}(\text{NO}_3)_2) = 207 + 16 \cdot 6 + 14 \cdot 2 = 331 \text{ г/моль}$$

$$15 W(\text{Pb}(\text{NO}_3)_2) = \frac{3,31}{150} \cdot 100\% = 2,2\%$$

$$\text{Ответ: } W(\text{Pb}(\text{NO}_3)_2) = 2,2\%$$

Умнож: 145